

## Read Book Guided Inquiry Balancing Chemical Equations

# Guided Inquiry Balancing Chemical Equations

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## **Balancing Chemical Equations Guided Inquiry Annotated**

BALANCING CHEMICAL EQUATIONS Load the simulation Balancing Chemical Equations CRITICAL THINKING QUESTIONS 1. Explore the Balancing Chemical Equations simulation. Discuss with your group what you find. a) What are the different ways that the simulation indicates when an equation is balanced? When arrows are yellow and there's a smiling face. 2. For each balanced reaction, indicate the total ...

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## **Balancing\_Chemical\_Equations\_Guided\_Inquiry\_StudentHandout ...**

Balancing Chemical Equations - Guided Inquiry Activity Description The learning goals of this guided-inquiry activity are as follows: Students will be able to: (1) Determine required conditions for a reaction to be considered "balanced" and relate these conditions to laws of matter, and (2) Develop

## **Guided Inquiry Balancing Chemical Equations**

Adapted from Chemistry: A Guided Inquiry, 2 nd ed. by Moog and Farrell (Wiley and Sons, Inc.) Balancing Chemical Equations PART I: Background Information: Word and Formula Equations The first step in writing a chemical equation is to identify the facts to be represented.

## **Balancing Chemical Equations - Welcome to nobel.scas.bcit.ca**

BALANCING CHEMICAL EQUATIONS 2 8. Start level 2 of the balancing equation game and adding new strategies as needed.

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9. In the simulation, were you able to use noninteger numbers (like  $\frac{1}{2}$  or 0.43) for the coefficients

## **BALANCING CHEMICAL EQUATIONS**

When you find difficulty in balancing the equation in the balancing chemical equations worksheet, you can miss it with a fraction of  $\frac{1}{2}$  and that will easily balance the equation. But the problem is that you cannot have a fraction for the co-efficient, this is why doubling all coefficients will help you balance the equation.

## **49 Balancing Chemical Equations Worksheets [with Answers]**

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## **Guided Inquiry Balancing Chemical Equations**

There are five basic types of chemical reactions (S, D, SR, DR, Comb.) that exhibit distinct patterns for the prediction of products. Experimental evidence can be used to identify products in a reaction. All reactions follow the Law of Conservation of Matter (balancing equations).

## **NOTES: CHEMICAL REACTIONS**

Displaying top 8 worksheets found for - Balancing Chemical Equations Answer Sheet. Some of the worksheets for this concept are Balancing equations practice problems, Teacher answer balancing equations, Balancing chemical equations answer, Chemical formulas equations work answers, Another balancing equation answer key, Balancing equations work answers, Balancing word equations chapter 9 ...

## **Balancing Chemical Equations Answer Sheet Worksheets**

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Balancing Chemical Equations - Guided Inquiry Activity. The

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learning goals of this guided-inquiry activity are as follows: Students will be able to: (1) Determine required conditions for a reaction to be considered “balanced” and relate these conditions to laws of matter, and (2) Develop strategies to balance chemical equations.

## **Balancing Chemical Equations - Guided Inquiry Activity ...**

Balancing Chemical Equations - Guided Inquiry Activity

Description The learning goals of this guided-inquiry activity are as follows: Students will be able to: (1) Determine required conditions for a reaction to be considered “balanced” and relate these conditions to laws of matter, and (2) Develop strategies to balance chemical equations.

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## **Balancing Chemical Equations Guided Inquiry StudentHandout-2**

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## **Balancing Chemical Equations - Chemical Equations ...**

Balance the following chemical equations. 1.  $\text{Fe} + \text{H}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + \text{H}_2$ : 2.  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{H}_2\text{O} + \text{CO}_2$ : 3.  $\text{KOH} + \text{H}_3\text{PO}_4 \rightarrow \text{K}_3\text{PO}_4 + \text{H}_2\text{O}$ : 4.  $\text{SnO}_2 + \text{H}_2 \rightarrow \text{Sn} + \text{H}_2\text{O}$ : 5.  $\text{NH}_3 + \text{O}_2 \rightarrow \text{NO} + \text{H}_2\text{O}$ : 6.  $\text{KNO}_3 + \text{H}_2\text{CO}_3 \rightarrow \text{K}_2\text{CO}_3 + \text{HNO}_3$ : 7.  $\text{B}_2\text{Br}_6 + \text{HNO}_3 \rightarrow \text{B}_2(\text{NO}_3)_3 + \text{HBr}$ : 8.  $3\text{Li} + \text{BF}_3 \rightarrow \text{Li}_2\text{SO}_3 + \text{B}_2(\text{SO}_3)_3 + \text{LiF}$ : 9.  $(\text{NH}_4)_3\text{PO}_4 + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{Pb} \dots$

## **Balancing Equations: Practice Problems**

Adapted from Chemistry: A Guided Inquiry, 2nd ed. by Moog and Farrell (Wiley and Sons, Inc.) Balancing Chemical Equations Teacher Notes Time Required: Estimated 20-30 minutes for Part I and 50 minutes for Part II.

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